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12 3 rate laws chemistry 2e openstax Apr 27 2020 remember that a number raised to the zero power is equal to 1 thus $x^0 = 1$

which is why the concentration term may be omitted from the rate law the rate of reaction is solely dependent on the concentration of NO_2 a later chapter section on reaction mechanisms will explain how a reactant's concentration can have no effect on a reaction rate despite being involved in the

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aqueous solution $[\text{OH}^-]$ in mol l⁻¹ can be calculated if the pOH of the solution is known pOH is defined as the negative logarithm to base 10 of the hydroxide ion concentration in mol l⁻¹ $\text{pOH} = \log_{10} [\text{OH}^-]$ $[\text{OH}^-]$ is the concentration of hydroxide ions in mol l⁻¹ mol l⁻¹ or $\text{m} [\text{OH}^-]$ in mol l⁻¹ can be calculated using the equation

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